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|  | AP Chemistry | Honors Chemistry |
| Monday | Christmas break | Christmas break |
| Tuesday | In-service | In-service |
| Wednesday | AP chem solutions  Unit 3 – Chemical Reactions I (Chapter   * Electrolytic Solutions (4.1) * Solubility (4.2) * Precipitation Reactions (4.2) * REDOX Reactions (4.4) * Oxidation Numbers (4.4)   HW: Aqueous Solutions and Chemical Reactions Worksheet I – Due tomorrow | 1. Discussion and Notes 9.1– Naming Ions.    1. Demo: Mg + O2 🡪 MgO    2. Demo: Gunpowder 2. Handout Ion Charge Sheet.   HW: Complete Section Review 9.1& 9.1 Interpreting Graphics |
| Thursday | Grade in class: HW: Aqueous Solutions and Chemical Reactions Worksheet I  Unit 3 – Chemical Reactions II (Slides #1 -14)   * **The Activity Series (4.4)**   + **Demo for slide #4**   + **Zn (s) + Cu 2- 🡪 Cu (s) + Zn 2-**   + **Cu (s) + Zn 2- 🡪 No Reaction** * **Halogen Displacement Reactions** * Hydrogen Displacement reactions (4.4) * Disproportionation Reactions * Combustion Reaction (3.2, 4.4)   HW: Aqueous Solutions and Chemical Reactions Worksheet II - #1 through #10 | 1. Grade: Section Review 9.1& 9.1 Interpreting Graphics. 2. 9.2 Naming & Writing Formulas for Ionic Compounds Through slide #39 |
| Friday | * SET: Draw Product beaker for reaction:   + **Zn (s) + Cu 2- 🡪 Cu (s) + Zn 2- and**   + **Write the half reactions and balnced net ionic equation for this reactions: A solid strip of Mg ribbon is dropped into silver nitrate solution**   Cont: Unit 3 – Chemical Reactions II (slides 15 -40)   * The Activity Series (4.4)   + Demo for slide #4   + Zn (s) + Cu 2- 🡪 Cu (s) + Zn 2-   + Cu (s) + Zn 2- 🡪 No Reaction * Halogen Displacement Reactions * **Hydrogen Displacement reactions (4.4)** * **Disproportionation Reactions** * **Combustion Reaction (3.2, 4.4)**   HW: Aqueous Solutions and Chemical Reactions Worksheet II - #11 through #26 | 1. Set: Practice write the chemical formulas of 15 Ionic compounds on the back of Lab Sheets. 2. Lab - 12: Chemical Names and Formulas 3. Review Lab #12, Chemical Names & Formulas |
|  | * What are we learning? AP Chemical Reactions * Why are we learning this? To build our basics for this semester * How will we know when we have learned this? By completing unit exam successfully | * What are we learning? Naming and writing chemical formulas * Why are we learning this? Necessary skill for semester success. * How will we know when we have learned this? Able to pass quiz over this material successfully. |

AP Chem Set for Friday:

* : Draw Product beaker for reaction:
  + **Zn (s) + Cu 2+ 🡪 Cu (s) + Zn 2+**

**Zn 2+**

**Zn 2+**

**Zn 2+**

**Zn 2+**

**Cu (s) 🡪**

* Write the half reactions and balanced net ionic equation for this reaction:
  + A solid strip of Mg ribbon is dropped into silver nitrate solution
    - **Oxidation half reaction: Mg (s) 🡪 Mg 2+ + 2 e-**
    - **Reduction half reaction: 2 Ag 1+ + 2 e-🡪 2 Ag (s)**
    - **Balanced Net Ionic equations: Mg (s) + 2 Ag 1+🡪 2 Ag (s) + Mg 2+**