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|  | AP Chemistry | Honors Chemistry |
| Monday | Christmas break | Christmas break |
| Tuesday | In-service | In-service |
| Wednesday | AP chem solutionsUnit 3 – Chemical Reactions I (Chapter * Electrolytic Solutions (4.1)
* Solubility (4.2)
* Precipitation Reactions (4.2)
* REDOX Reactions (4.4)
* Oxidation Numbers (4.4)

HW: Aqueous Solutions and Chemical Reactions Worksheet I – Due tomorrow | 1. Discussion and Notes 9.1– Naming Ions.
	1. Demo: Mg + O2 🡪 MgO
	2. Demo: Gunpowder
2. Handout Ion Charge Sheet.

HW: Complete Section Review 9.1& 9.1 Interpreting Graphics |
| Thursday | Grade in class: HW: Aqueous Solutions and Chemical Reactions Worksheet IUnit 3 – Chemical Reactions II (Slides #1 -14)* **The Activity Series (4.4)**
	+ **Demo for slide #4**
	+ **Zn (s) + Cu 2- 🡪 Cu (s) + Zn 2-**
	+ **Cu (s) + Zn 2- 🡪 No Reaction**
* **Halogen Displacement Reactions**
* Hydrogen Displacement reactions (4.4)
* Disproportionation Reactions
* Combustion Reaction (3.2, 4.4)

HW: Aqueous Solutions and Chemical Reactions Worksheet II - #1 through #10 | 1. Grade: Section Review 9.1& 9.1 Interpreting Graphics.
2. 9.2 Naming & Writing Formulas for Ionic Compounds Through slide #39
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| Friday | * SET: Draw Product beaker for reaction:
	+ **Zn (s) + Cu 2- 🡪 Cu (s) + Zn 2- and**
	+ **Write the half reactions and balnced net ionic equation for this reactions: A solid strip of Mg ribbon is dropped into silver nitrate solution**

 Cont: Unit 3 – Chemical Reactions II (slides 15 -40)* The Activity Series (4.4)
	+ Demo for slide #4
	+ Zn (s) + Cu 2- 🡪 Cu (s) + Zn 2-
	+ Cu (s) + Zn 2- 🡪 No Reaction
* Halogen Displacement Reactions
* **Hydrogen Displacement reactions (4.4)**
* **Disproportionation Reactions**
* **Combustion Reaction (3.2, 4.4)**

HW: Aqueous Solutions and Chemical Reactions Worksheet II - #11 through #26 | 1. Set: Practice write the chemical formulas of 15 Ionic compounds on the back of Lab Sheets.
2. Lab - 12: Chemical Names and Formulas
3. Review Lab #12, Chemical Names & Formulas
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|  | * What are we learning? AP Chemical Reactions
* Why are we learning this? To build our basics for this semester
* How will we know when we have learned this? By completing unit exam successfully
 | * What are we learning? Naming and writing chemical formulas
* Why are we learning this? Necessary skill for semester success.
* How will we know when we have learned this? Able to pass quiz over this material successfully.
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AP Chem Set for Friday:

* : Draw Product beaker for reaction:
	+ **Zn (s) + Cu 2+ 🡪 Cu (s) + Zn 2+**

**Zn 2+**

**Zn 2+**

**Zn 2+**

**Zn 2+**

**Cu (s) 🡪**

* Write the half reactions and balanced net ionic equation for this reaction:
	+ A solid strip of Mg ribbon is dropped into silver nitrate solution
		- **Oxidation half reaction: Mg (s) 🡪 Mg 2+ + 2 e-**
		- **Reduction half reaction: 2 Ag 1+ + 2 e-🡪 2 Ag (s)**
		- **Balanced Net Ionic equations: Mg (s) + 2 Ag 1+🡪 2 Ag (s) + Mg 2+**