**Chap 10 Honors Chemistry**

**TEST REVIEW**

**Matching**

Define the following

1. molar volume
2. molar mass
3. atomic mass
4. representative particle
5. percent composition
6. mole
7. standard temperature and pressure
8. Avogadro's number
9. empirical formula

**Remember & be able to:**

* *The mole is the SI unit is used to measure the number of representative particles in a substance*
* *Be able to list the seven diatomic molecules.*
* Avogadro's number of representative particles is equal to one mole
  + **Be able to work problems #51 through 61 on p. 338 in the textbook**
* *The atomic masses of any two elements contain the same number of atoms.*
* Be able to solve the following types of chemistry measurement problems:
  + Mole – Mass (Vice-Versa)
  + Mass - Mole –Volume (Vice-Versa)
  + Mass - Mole - # Particles (Vice-Versa)
  + Mole - # Particles (Vice-Versa)
  + Be able to work problems #62 through 65 on p. 338 in the textbook
* *The volume of one mole of a substance is 22.4 L at STP for all gases*
* *The molar volume of a gas at STP occupies 22.4 L*
* *Standard temperature and pressure, STP = 0°C & 101.3 kPa*
* *The molar mass of a gas can be determined from the density of the gas at STP*
  + **Be able to work Problems #66 through 70 on p. 338 in the textbook**
* Be able to:
  + calculate the percent composition of a compound
  + calculate the empirical formula of a compound
  + calculated the molecular formula of a compound.
  + **Be able to work problems #71 through 79 on p. 338-339 in the textbook**





